

## Notes

- 1 Tannenberger [33-tan] reported data for two samples of diacetylene. The first sample was prepared by pyrolysis, the second one by polymerization of acetylene. The data for the second sample are in good agreement with that reported by Strauss and Kollek [26-strkol] while data for the second sample obtained by pyrolysis are higher by about 10%. Therefore data from [26-strkol] and that of [33-tan] observed for the sample prepared by the polymerization of acetylene were selected for the evaluation of Antoine constants.
- 2 The source does not contain original experimental data. The rating provided is based on auxiliary information.
- 3 The source does not contain original experimental data. The rating provided and applicable temperature range estimates are based on auxiliary information.
- 4 Antoine constants have been obtained by the modification of more complicated equation used in the original source.
- 5 The data in the temperature range 338 K to 384 K is related to undercooled liquid.
- 6 The data by Ambrose *et al.* [76-amblaw] was selected. The data reported by Colomina *et al.* [89-coljim] is lower of about 10% and that by Overberger *et al.* [69-oveste] is higher of about 3% than the data selected.
- 7 Data reported by Osborne *et al.* [75-osbdou] was selected for the temperature interval of 348 K to 388 K while that by Inomara *et al.* [86-inoara-1] was selected for the temperature interval of 307 K to 348 K. Note that data by Osborne *et al.* is in good agreement with that of Sasse *et al.* [88-sasjos].
- 8 Kana'an [72-kan] reported experimental data obtained by both Knudsen and torsion methods. The data obtained by the torsion method was selected. The data obtained by Knudsen method is about 17% lower.
- 9 The Antoine constants were determined on the basis of two experimental data points only available. It is related to the fact that the melting point is close to 253 K, whereas the compound may explode at the temperatures higher than 293 K. Therefore the rating provided is a rough estimate only.
- 10 The difference between normal boiling point calculated from the equation reported in the original source and that corresponding to the equation with the parameters reported here is less than 5% (rating C).
- 11 The difference between normal boiling point calculated from the equation reported in the original source and that corresponding to the equation with the parameters reported here is greater than 5% (rating D).
- 12 The Antoine parameters were determined on the basis of two experimental points.
- 13 The Antoine parameters were determined on the basis of three experimental points.
- 14 The intersection point of the vapor pressure curves for the solid and for the liquid phases is  $T = 229.5\text{K}$ ,  $P = 0.0239\text{ kPa}$ .
- 15 The Antoine parameters in the temperature range 233 to 264 K were calculated based original experimental data [63-quinow] and Cox-Othmer diagram. Partial decomposition of the compound is reported at temperatures above 400 K.
- 16 The parameter A reported in the original source [66-banhas] is erroneous. We have corrected it to make it consistent with the reported value of the boiling point.
- 17 The values of the parameters are those reported in the original source. We have estimated the rating for the temperature range below 292 K.
- 18 Only the mean temperature of the temperature interval studied was reported in the original source.
- 19 There are two sources of original experimental data for this compound which are in good agreement with each other (the difference over the temperature interval is about 1.5%). We have presented three different sets of the Antoine parameters (one for each data set reported and one for the combination of all the data).

- 
- 20 The constant A originally reported was changed.
- 21 These parameters are based on the most recent data available for the compound.
- 22 The parameters are obtained by the modification of the equation reported in the original source.
- 23 The intersection point of the vapor pressure curves for the solid and for the liquid phases is  $T = 303.8\text{K}$ ,  $P = 0.0429\text{ kPa}$  (melting point is  $304.502\text{ K}$  [54-fin/gro] ).
- 24 The intersection point of the vapor pressure curves for the solid and for the liquid phases is  $T = 375.8\text{K}$ ,  $P = 0.0429\text{ kPa}$ .
- 25 The intersection point of the vapor pressure curves for the solid and for the liquid phases is  $T = 432.4\text{K}$ ,  $P = 0.2256\text{ kPa}$ .
- 26 The original source does not contain experimental data. The parameters of the equation were forwarded from the original source.
- 27 The parameters of the equation are forwarded from the original source. The information about the temperature range 'covered' by this equation is not provided in the original source.
- 28 The Antoine constants reported [72-kunwai] are related to the supercooled liquid.
- 29 The equation was developed based on single data points reported in various publications.
- 30 The source does not contain original experimental data. The parameters of the equation (up to  $354\text{ K}$ ) are forwarded from the original source. The rating provided is based on auxiliary information.
- 31 The applicable temperature range is determined from the graphical representation reported in the original source.
- 32 The equation reported in the original source is adjusted to ITS-90. The reported experimental value of normal boiling point is higher than the calculated one by  $1.5\text{K}$ .
- 33 The equation for the liquid phase (up to  $178\text{ K}$ ) reported in the original source is forwarded. No experimental data is reported in the original source.
- 34 The equation reported in the original source is adjusted to ITS-90.
- 35 Melting point of the commercial sample used is  $327\text{ K} < T_b < 329\text{ K}$ .
- 36 Only those polychlorobiphenyls are evaluated for which experimental vapor pressure data were reported at least at two different temperatures.
- 37 The recommended equation is the result of the fit of combination of the experimental data obtained by two different methods.
- 38 The equation reported in the original source is forwarded.