

Notes

1. The source does not contain original experimental data. The rating provided and the applicable temperature range estimates are based on auxiliary information.
2. The source does not contain original experimental data. The rating provided is based on auxiliary information.
3. The source does not contain original experimental data. The equation is formulated for the liquid phase. The rating provided is based on auxiliary information.
4. The source provides the Antoine parameters for the liquid phase. The rating and the applicable range are based on auxiliary information.
5. The Antoine parameters are calculated based on 2 experimental points.
6. The Antoine parameters are calculated based on 3 experimental points.
7. The source does not contain original experimental data in the numerical form. The experimental data are represented in the graphical form and in the form of the equation. The applicable range is determined from the graph. The rating provided is based on auxiliary information.
8. The Antoine parameters are calculated based on the graphical representation of the original experimental data.
9. The Antoine parameters are calculated based on 1 numerical data point and a slope of the function $\log P$ vs. $1/T$ expressed graphically.
10. The Antoine parameters correspond to the equation provided in the source with the adjustment to the temperature scale ITS-90.
11. The Antoine parameters are found using modification of the originally reported more complicated equation.
12. The source contains the original experimental data obtained by both dynamic and static methods. The only data used to evaluate the Antoine constants are those obtained by the static method since these data are more accurate in accordance to the original source.
13. The source reports the values of melting (187.4 K to 187.7 K) and normal boiling (314 K) points. Therefore the applicable range accepted is 187 K to 314 K.
14. The source provides information for the vapor – liquid equilibria as well as normal boiling point (338 K). We estimate the applicable temperature range from 273 K to 338 K.
15. The source reports the equation developed based on the original data published in [1895-ano].
16. The equation is developed using 2 single data points from 2 different sources.
17. No original experimental data is given in the source. The rating provided is our estimate.
18. The data reported in [60-davtho] are about 20% lower than those corresponding to the selected equation.
19. The data reported in various sources for the compound differ substantially. That probably reflects upon the fact that the sample for the compound might undergo chemical changes during evaporation.
20. The original source provides 6 data points having considerable dispersion on the fit $\log P = f(1/T)$. In order to minimize the dispersion (to 4.9%) we exclude the point of 3 Torr at 324.55 K.
21. The parameters of the Antoine equation are derived using the boiling point at the pressure of 10 Torr and the slope of the curve $\log P = f(1/T)$ reported in the original source.
22. The significant difference of the data reported by various authors might be caused by the fact that the compound can exist in 2 different forms (keto- and enol-) which may or may not be in equilibrium.
23. The parameters of the equation are obtained from the graphical representation of the $\log P = f(1/T)$ dependence reported in [48-feifis].
24. The compound undergoes decomposition at the temperatures above 398 K.
25. The parameters of the Antoine equation are derived using the 1 of 2 data points reported in [48-feifis] (10 Torr at 384 K) and the slope of the curve $\log P = f(1/T)$ reported in the same source.
26. The parameters of the Antoine equation are obtained as a fit for 2 data sets measured by different methods.
27. The original source reports experimentally obtained data in the range 318 K to 453 K in the form of the equation and the tabulated data extrapolated to 293 K. The rating provided is our estimate.
28. The parameters of the Antoine equation are those reported in [68-khobyk].

29. The equation is related to the undercooled liquid in the temperature range of 451 K to 452.6 K.
30. The source reports the vapor pressure over the commercial sample. It is commonly assumed that this type of the sample might be a mixture of 3,7-dimethyl-2,7-octadiene-1-ol and 3,7-dimethyl-2,6-octadiene-1-ol.
31. The parameters of the Antoine equation are obtained using data obtained by both static (pressure gauge) and dynamic (air as a carrier gas) methods.
32. The 'cr₁' designates rhombic form, 'cr₂' is related to the monoclinic form.
33. The selected parameters of the equation are based on a combination of the experimental data from [49-sausta] and other sources.
34. The selected parameters of the equation are based on a combination of the tabulated [56-ano-17] and graphically displayed [77-mel] data.
35. The selected parameters of the Antoine equation are based on 4 of 6 experimental points reported in the original source. The other 2 points (23.8 Torr and 76.2 Torr) are not fitted well by the Antoine function.
36. The selected equation is related to undercooled liquid. The melting point reported in [44-parwei] is 300 K.
37. The original source reports a normal boiling point and a vapor pressure equation for the liquid phase. The rating is our estimate.
38. The selected equation is based on the data reported in the graphical form.
39. The parameters of the Antoine equation are determined using 2 data points and the graphical form of the vapor pressure/temperature dependence.
40. There is a substantial difference in the experimental data published in various sources. For example, the data reported in [40-hierei] differ from those published in [47-stu] by more than 10%. The selected parameters of the Antoine equation are based on experimental data taken from [77-matmun]. This equation is consistent closely with the data reported in [40-hierei] and [58-klo-1].
41. The parameters of the Antoine equation were derived from those of Chebyshev polynomial series reported in [76-ambell].
42. The equation is related to undercooled liquid at the temperatures below 300 K.
43. The parameters of the Antoine equation are those reported in the original publication.
44. The equation is developed using 2 tabulated data points and the graphical form of the vapor pressure/temperature dependence.
45. The equation is related to undercooled liquid at the temperatures below 317.65 K.
46. The compound decomposes at the temperatures above 445 K.
47. The critical temperature is 500.17 K.
48. The parameters of the Antoine equation are determined using data provided in a private communication.
49. The equation is based on the experimental data determined by the effusion method.
50. The rating provided is our estimate.
51. The parameters of the Antoine equation are derived based on the average of the original experimental data obtained with the use of 2 apparatuses at every temperature.
52. The equation is related to undercooled liquid at the temperatures below 300 K.
53. Tri-*tert*-butylmethanol undergoes a solid phase transition of 'rigid' crystal-to-plastic crystal type at 302.17 K. At 390.15 K a plastic crystalline phase has a melting point. The selected equation is related to metastable undercooled plastic crystal phase in the temperature range 278 K to 302 K. Above 302 K the equation is describing stable plastic crystalline phase.
54. The data reported in [66-grabur] are considerably higher than those fitted to the selected equation.
55. The selected equation is taken from the original source. The temperature range corresponds to the one used in the graphical presentation of the data in the original source. The rating provided is our estimate.
56. The equation reported in [58-hoypep] yields the vapor pressure data substantially higher than those corresponding to the selected equation.
57. Some data used to develop the equation are obtained by the extrapolation of the experimental values reported in the original source.
58. The rating provided is our estimate.
59. The compound decomposes at the temperatures higher than 690 K.
60. The original source provides the selected equation for the commercial sample at the temperatures below the melting point. The rating is our estimate.

61. The selected equation is taken from the original source. The rating provided is our estimate.
62. There is a great number of sources reporting the vapor pressure data for the compound. However even after preliminary selection of those having the least uncertainties, data are still very scattered which does not allow to achieve a quality rating higher than "C" and "D".
63. The data reported in [87-daujal] are about 4% higher than those given in [81-ambhal].
64. The compound decomposes at the temperatures higher than the normal boiling point.
65. The compound is in the state of undercooled liquid at the temperatures below 365 K.
66. The original equation reported in the source is modified. The range and the rating are our estimates.
67. The Antoine equation is developed using a combination of the tabulated and graphical data.
68. The sample described in the source is of very poor quality. It contains uncharacterized amount of the volatile impurities.
69. To develop the selected equation some data at the temperatures higher than 430 K are taken from the Wagner equation fit provided in the original source.